

Teacher Name: Lillie  
Class: CP (NGSS) Chem  
Period: 1, 2, 3, 6  
Assignment: Assignment Week 2  
**Due: Friday, 5/8**

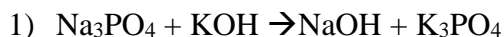
Student Name: \_\_\_\_\_

### Types of Chemical Reactions

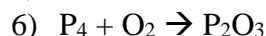
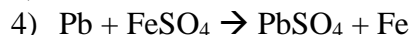
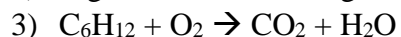
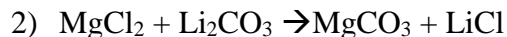
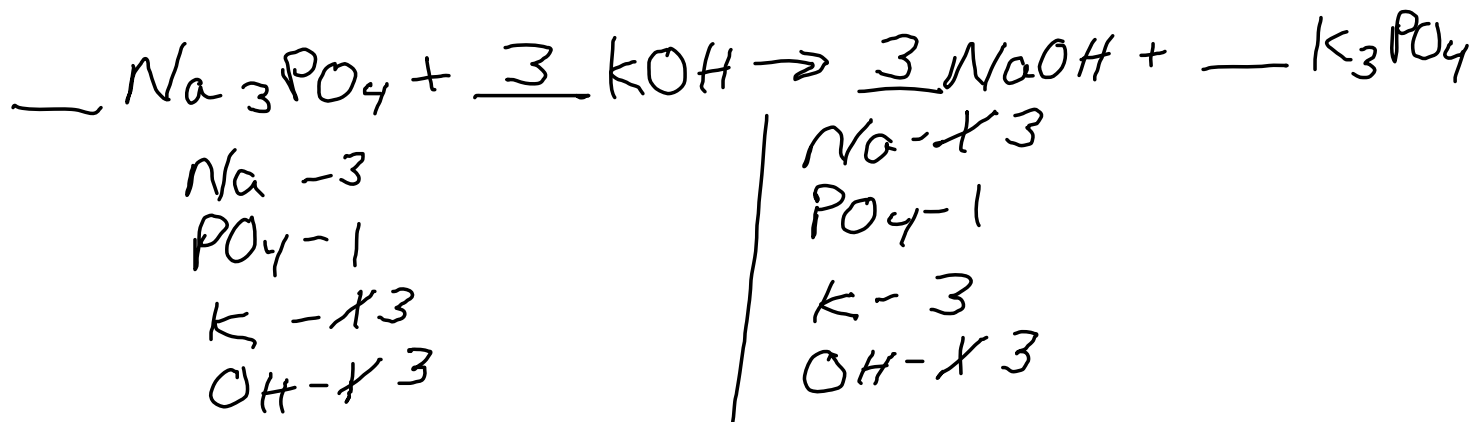
- Skills:
  - ✓ identify the 5 general types of reactions
  - ✓ reinforce balancing equations
  - ✓ complete and balance chemical equations
- General instructions:
  - ✓ Read section 11.2
  - ✓ Complete practice problems and section assessment from monad
  - ✓ Complete the problems from Tuesday through Friday
- Questions:
  - ✓ Please send email or attend online office hours if you have any questions

Date	Activity ( <b>Work to submit</b> )
Monday (4/27)	Read Section 11.2 Take notes ( <b>submitted</b> ) Do practice problems 13,14,15,16,17,18,19,20,21 ( <b>submitted</b> ) Do section assessment 24,25,26,27 ( <b>submitted</b> )
Tuesday (4/28)	Do problems 1-6 (below) (note that #1 is already completed for you as an example. <b>Use this set up for all of your problems</b> ) ( <b>submitted</b> )
Wednesday (4/29)	Do problems 7-11 ( <b>submitted</b> )
Thursday (4/30)	Do problems 12-17 (below) (note that #12 is already completed for you as an example. <b>Use this set up for all of your problems</b> ) ( <b>submitted</b> )
Friday (5/1)	Do problems 18-22 <b>NOTE: 23-28 are for you do get more practice if you want.</b> ( <b>submitted</b> )

### Identify the Type of reaction and balance the equation



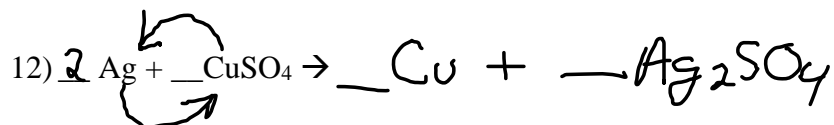
Type of reaction  
Double Replacement



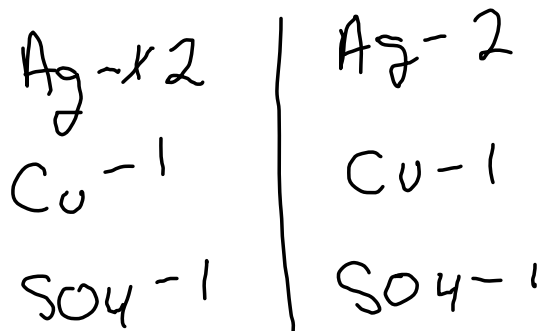
- 7)  $\text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu}(\text{NO}_3)_2 + \text{Ag}$   
 8)  $\text{C}_3\text{H}_6\text{O} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$   
 9)  $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + \text{Cl}_2$   
 10)  $\text{O}_3 \rightarrow \text{O} + \text{O}_2$   
 11)  $\text{NO}_2 \rightarrow \text{O}_2 + \text{N}_2$

Identify the reaction, Complete the equation and Balance the equation

Type of Reaction



Single Replacement



- 13)  $\text{NaI} + \text{CaCl}_2 \rightarrow$   
 14)  $\text{O}_2 + \text{H}_2 \rightarrow$   
 15)  $\text{HNO}_3 + \text{Mn}(\text{OH})_2 \rightarrow$   
 16)  $\text{HCN} + \text{CuSO}_4 \rightarrow$   
 17)  $\text{H}_2\text{O} + \text{AgI} \rightarrow$   
 18)  $\text{LiBr} + \text{Co}(\text{SO}_3)_2 \rightarrow$   
 19)  $\text{LiNO}_3 + \text{Ag} \rightarrow$   
 20)  $\text{AlCl}_3 + \text{Cs} \rightarrow$   
 21)  $\text{Al}(\text{NO}_3)_3 + \text{Ga} \rightarrow$   
 22)  $\text{H}_2\text{SO}_4 + \text{NH}_4\text{OH} \rightarrow$   
 23)  $\text{CH}_3\text{COOH} + \text{O}_2 \rightarrow$   
 24)  $\text{C}_4\text{H}_8 + \text{O}_2 \rightarrow$   
 25)  $\text{KCl} + \text{Mg}(\text{OH})_2 \rightarrow$   
 26)  $\text{Zn} + \text{Au}(\text{NO}_2)_2 \rightarrow$   
 27)  $\text{BaS} + \text{PtCl}_2 \rightarrow$   
 28)  $\text{Na}_2\text{O} \rightarrow$