Teacher Name: Lillie Student Name: _____
Class: CP (NGSS) Chem
Period: 1, 2, 3, 6
Assignment: Assignment Week 2
Due: ***Friday, 5/8***

Types of Chemical Reactions

- Skills:
  - ✓ identify the 5 general types of reactions
  - ✓ reinforce balancing equations
  - ✓ complete and balance chemical equations
- General instructions:
  - ✓ Read section 11.2
  - ✓ Complete practice problems and section assessment from monad
  - ✓ Complete the problems from Tuesday through Friday
- Questions:
  - ✓ Please send email or attend online office hours if you have any questions

Date	Activity (**Work to submit**)
Monday (4/27)	Read Section 11.2 Take notes **(submitted)** Do practice problems 13,14,15,16,17,18,19,20,21 **(submitted)** Do section assessment 24,25,26,27 **(submitted)**
Tuesday (4/28)	Do problems 1-6 (below) (note that #1 is already completed for you as an example. **Use this set up for all of your problems**) **(submitted)**
Wednesday (4/29)	Do problems 7-11 **(submitted)**
Thursday (4/30)	Do problems 12-17 (below) (note that #12 is already completed for you as an example. **Use this set up for all of your problems**) **(submitted)**
Friday (5/1)	Do problems 18-22 ***NOTE: 23-28 are for you do get more practice if you want***. **(submitted)**

Identify the Type of reaction and balance the equation

1) $Na_3PO_4 + KOH \rightarrow NaOH + K_3PO_4$

Type of reaction
Double Replacement

$\_\ Na_3PO_4 + \underline{3}\ KOH \rightarrow \underline{3}\ NaOH + \_\ K_3PO_4$

Na -3
PO_4 - 1
K -X3
OH-X3

Na-X3
PO_4-1
K-3
OH-X3

2) $MgCl_2 + Li_2CO_3 \rightarrow MgCO_3 + LiCl$
3) $C_6H_{12} + O_2 \rightarrow CO_2 + H_2O$
4) $Pb + FeSO_4 \rightarrow PbSO_4 + Fe$
5) $CaCO_3 \rightarrow CaO + CO_2$
6) $P_4 + O_2 \rightarrow P_2O_3$

7) $AgNO_3 + Cu \rightarrow Cu(NO_3)_2 + Ag$
8) $C_3H_6O + O_2 \rightarrow CO_2 + H_2O$
9) $SeCl_6 + O_2 \rightarrow SeO_2 + Cl_2$
10) $O_3 \rightarrow O + O_2$
11) $NO_2 \rightarrow O_2 + N_2$

Identify the reaction, Complete the equation and Balance the equation

Type of Reaction: Single Replacement

12) 2 Ag + $CuSO_4$ → Cu + Ag_2SO_4

Ag - x2	Ag - 2
Cu - 1	Cu - 1
SO4 - 1	SO4 - 1

13) __ NaI + __ $CaCl_2$ →
14) __ O_2 + __ H_2 →
15) __ HNO_3 + __ $Mn(OH)_2$ →
16) __ HCN + __ $CuSO_4$ →
17) __ H_2O + __ AgI →
18) __ LiBr + __ $Co(SO_3)_2$ →
19) __ $LiNO_3$ + __Ag →
20) __ $AlCl_3$ + __ Cs →
21) __ $Al(NO_3)_3$ + __ Ga →
22) __ H_2SO_4 + __ NH_4OH →
23) __ CH_3COOH + __ O_2 →
24) __ C_4H_8 + __ O_2 →
25) __ KCl + __ $Mg(OH)_2$ →
26) __ Zn + __ $Au(NO2)_2$ →
27) __ BaS + __ $PtCl_2$ →
28) __ Na_2O →